

Effects of sodium chloride on the properties of chlorophyll *a* submonolayer adsorbed onto hydrophobic and hydrophilic surfaces using broadband spectroscopy with single-mode integrated optical waveguides

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Abstract. In this work, we experimentally investigated the effects of sodium chloride on the molar absorptivity and surface density of a submonolayer of chlorophyll *a* adsorbed onto hydrophilic and hydrophobic solid/liquid interfaces. Those investigations were made possible by a broadband spectroscopic platform based on single-mode, integrated optical waveguides, which allows for extremely sensitive spectroscopic detection of analytes immobilized at submonolayer levels. Chlorophyll *a* with a constant bulk concentration (1.4 μM) was dissolved in phosphate buffer solutions (7 mM) of neutral pH, but with different sodium chloride concentrations. For a buffer solution of 1 mM of sodium chloride, the measured surface density of chlorophyll *a* was 0.209 pmol/cm² for a hydrophilic and 0.125 pmol/cm² for a hydrophobic surface. For a phosphate buffer solution of 10 mM of sodium chloride, the measured surface density of chlorophyll *a* was 0.528 pmol/cm² for a hydrophilic and 0.337 pmol/cm² for a hydrophobic surface. Additionally, a hypsochromic shift of the Soret band was observed for the adsorbed pigment in correlation with an increase in buffer ionic strength. The adsorption of chlorophyll *a* onto different surfaces can play an important role to elucidate several processes found in nature and provide a rationale for bio-inspired new material technologies.
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Subject terms: chlorophyll *a*; molecular surface adsorption; waveguide spectroscopy; atomic layer deposition; molecular vapor deposition silane; perfluorodecyltrichlorosilane.

Paper 100893SSR received Oct. 30, 2010; revised manuscript received Feb. 10, 2011; accepted for publication Feb. 15, 2011; published online Apr. 21, 2011.

1 Introduction

Chlorophyll *a* is a key biochemical component in the molecular apparatus that is responsible for photosynthesis, the critical process in which the energy from sunlight is used to produce life-sustaining oxygen. In the photosynthetic reaction sequence, water is oxidized to oxygen^{1,2} and chlorophyll powers this redox chemistry both by absorbing light to attain an excited electronic state and also by transferring this excitation energy to the photosystem reaction centers, where the energy is trapped in a photochemical reaction.^{3–5} Most of the chlorophyll molecules found in nature are bound to a class of photosynthetic proteins, called light harvesting complex,⁶ embedded in the thylakoid membranes of chloroplasts present in plants and algae cells.⁷ The principal chromophore of a chlorophyll *a* molecule is a porphyrin ring, coordinated to a central magnesium metal atom (see Fig. 1). Chlorophyll *a* absorbs light in the visible spectrum at approximately 400 to 440 nm (Soret band) and 660 to 680 nm (Q band). The peak position in both regions strongly depends on the solvent composition surrounding the pigment.^{8,9} Chlorophyll *a* also fluoresces at the 680 to 688 nm spectral region when excited by light with a shorter wavelength (e.g., 435 nm).¹⁰ Those

optical properties from chlorophyll *a* are used for several applications such as the identification of microalgae and phytoplankton in the ocean,^{11,12} the study of reactions occurring at different time scales in plants and algae cells,^{13–15} and recently have been investigated for performance improvement of dye-sensitized thin-film solar cells.^{16,17} In most of these applications, chlorophyll *a* is found in contact with a surface (e.g., cell membranes, thin-film solid surface); however, due to the lack of sensitivity in conventional instrumentation, there are only a few reports in the literature characterizing chlorophyll *a* adsorbed to surfaces at monolayer or submonolayer levels,^{18–21} and none of them consider the combined effect of the solution composition and the surface chemistry on the optical properties of the adsorbed pigment. When a liquid solution comes in contact with a solid surface, several competing forces can eventually drive the surface adsorption of molecules that were originally dissolved in the solution.²² Surface-adsorbed biomolecules will then be exposed to a fairly different nano-environment that may eventually alter the molecular optical properties.²³ The composition of the solvent that is used to dissolve a biomolecule also plays an important role in the adsorption process by altering the affinity of the molecule with the surface.²⁴

Single-mode integrated optical waveguides is a highly sensitive technique for investigation of surface-adsorbed

cases, the temperature of the precursors was kept at 55°C and the reaction chamber temperature was 35°C. Before the deposition of the organic films, a 60-s oxygen-plasma cleaning was performed to remove any organic residue from the alumina surface. The presence of the aluminum oxide layer, compared to metal or polymeric surfaces, improves the stability and durability of the organic films deposited because of the higher density of the reactive hydroxyl groups on the surface of the oxide film.³⁹

The chlorophyll *a* solution was injected into the flowcell with a volume of approximately 2 ml covering the waveguide surface. All the data acquisition was performed at room temperature (25°C).

2.3 Absorbance Spectra of Chlorophyll *a* in Solution and Surface

The concentration of chlorophyll *a* in solution was obtained by measuring the absorbance at the Q band (672 nm) in solution with a conventional spectrophotometer (Cary 300, Varian). The molar absorptivity of chlorophyll *a* in solution is well known ($\epsilon_{672} = 73,265 \text{ M}^{-1} \text{ cm}^{-1}$).⁴⁰ To ensure that the chlorophyll *a* concentration remained unaltered during our experiment, the sample to be measured was collected after the passage through the waveguide flowcell. Figure 5 shows the absorbance spectra obtained for the solutions with the same buffer pH and chlorophyll *a* concentration, but different concentrations of NaCl; as expected they exhibit a good agreement with each other.

To obtain the absorbance spectra of chlorophyll *a* that were surface-adsorbed onto the functionalized waveguide surface, the integration time for data acquisition in the CCD was fixed at 60 s. The sequence of data collection was the following: first, the buffer solution (without chlorophyll *a*) was injected into the flowcell and a reference signal was acquired; next, the chlorophyll *a* solution was introduced into the flowcell and data was sequentially acquired every 10 min until equilibrium in the amount of surface-adsorbed chlorophyll *a* was reached. A dark current signal was finally acquired by blocking the input light beam.

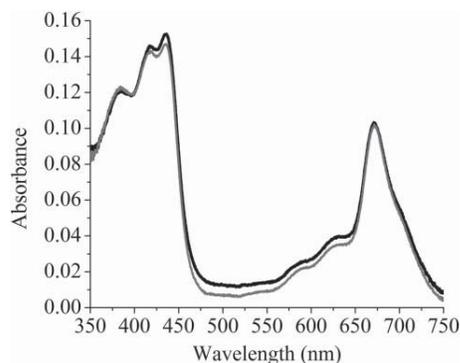


Fig. 5 Absorbance spectra for 1.4 μM of chlorophyll in solutions of different ionic strength. The black curve is chlorophyll *a* dissolved in 7 mM sodium phosphate plus 10 mM of sodium chloride, pH 7.25. The gray curve is chlorophyll *a* dissolved in 7 mM sodium phosphate plus 1 mM of sodium chloride, pH 7.25. The different concentration of sodium chloride does not markedly affect the absorbance of chlorophyll *a* in solution.

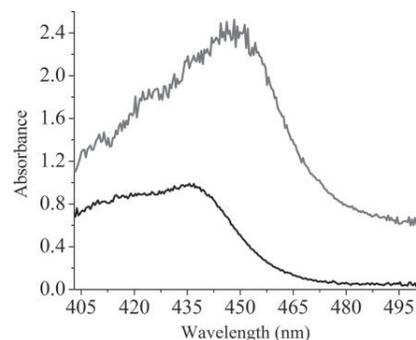


Fig. 6 Absorbance spectra of chlorophyll *a* adsorbed to a hydrophilic surface, acquired after adsorption equilibrium was reached. In both spectra, the amount of chlorophyll *a* dissolved in solution was the same (1.4 μM), as was the pH of the 7 mM sodium phosphate buffer (7.25). Black shows that the concentration of NaCl was 1 mM; gray shows that the concentration of NaCl was 10 mM.

3 Results and Discussions

3.1 Absorbance Measurements of Chlorophyll *a* Adsorbed to Chemically Modified Waveguide Surfaces

The absorbance data for surface-adsorbed chlorophyll *a* are presented in Fig. 6 for the aluminum oxide waveguide functionalized with a hydrophilic surface, and in Fig. 7 for the hydrophobic surface. In all cases, the solution concentration of chlorophyll *a* injected into the flowcell was held constant, but the sodium chloride concentration was varied at two different levels: 1 and 10 mM. The spectra presented in Figs. 6 and 7 correspond to the situation where equilibrium had been reached for the surface-adsorbed species, i.e., enough time had elapsed that no further changes in the spectra were observed. From these data we qualitatively observed that chlorophyll *a* adsorbs to a higher degree to the hydrophilic surface. In order to quantitatively describe the surface density, the sensitivity of the waveguide spectrometer had to be considered.

3.2 Sensitivity of the Waveguide Spectrometer

The sensitivity factor, S , which is defined as the ratio between the absorbance measured using a waveguide platform, A_{WG} ,

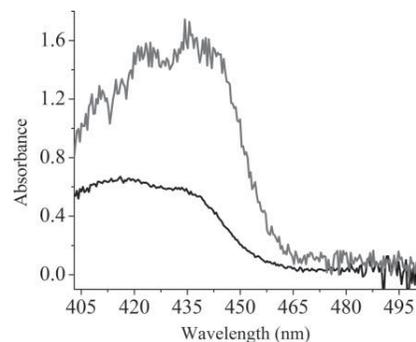


Fig. 7 Spectra of chlorophyll *a* adsorbed to a hydrophobic functionalized surface, acquired after the steady state was reached. In both spectra, the amount of chlorophyll *a* dissolved in solution was the same (1.4 μM), as was the pH of the 7 mM sodium phosphate buffer (7.25). Black shows that the concentration of NaCl was 1 mM; gray shows that the concentration of NaCl was 10 mM.

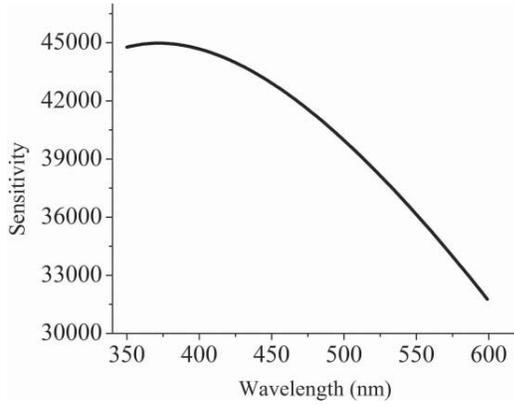


Fig. 8 Sensitivity values calculated for a single-mode waveguide with 178 nm of thickness and considering the following values for the constants: $L = 3.4$ cm, $n_c = 1.33$, with the approximation that $n_l \approx n_c$, $n_s = 1.51$. The effective refractive index, N_{TE} , and the effective thickness, $t_{eff,TE}$, were calculated from the waveguide dispersion equation. The refractive index of the alumina film was calculated using the relation $n_w(\lambda) = a + b/\lambda^2 + c/\lambda^4$, where $a = 1.645\ 76$, $b = 42.898\ 98$, and $c = 308\ 958\ 233.142$ from Ref. 32.

and the absorbance measured in direct transmission, A_{TR} , for an arbitrary layer of chromophores in both cases, is given by^{26,30}

$$S = \frac{A_{WG}}{A_{TR}} = \frac{A_{WG}}{\epsilon_{surf}\Gamma}, \quad (1)$$

where ϵ_{surf} is the molar absorptivity of the surface-bound chromophore and Γ is the chromophore surface density. The sensitivity factor is polarization-dependent and is given by the following expression for the TE-polarization:²⁶

$$S_{TE} = \frac{2n_l(n_w^2 - N_{TE}^2)L}{N_{TE}(n_w^2 - n_c^2)t_{eff,TE}}, \quad (2)$$

where n_l is the real part of the refractive index of the adsorbed analyte, n_w is the refractive index of the alumina waveguiding film, n_c is the refractive index of the buffer solution, N_{TE} is the effective refractive index of the waveguide, $t_{eff,TE}$ is the effective thickness of the waveguide (which is related to physical thickness of the guiding film, t), and L is the distance between the input and output couplers. For the single-mode waveguide used in this work, the sensitivity factor is presented in Fig. 8 as a function of the wavelength for the spectral region of interest in this work. It is the extremely high value of the sensitive factor that makes possible the investigation of very low amounts of surface adsorbed species.

3.3 Determination of Surface Coverage and Kinetics of Chlorophyll a Adsorbed to Different Functionalized Surfaces

In order to calculate the surface density of chlorophyll a , we solved Eq. (1) for Γ

$$\Gamma = \frac{A_{WG,TE}}{S_{TE}\epsilon_{surf}}, \quad (3)$$

and used the experimental results of the peak absorbance, $A_{WG,TE}$, measured by our single-mode waveguide-based spectrometer and the results of the sensitivity factor, S_{TE} , from Eq. (2). In Fig. 9 we display the surface-adsorption kinetic results of chlorophyll a (surface density of chloro-

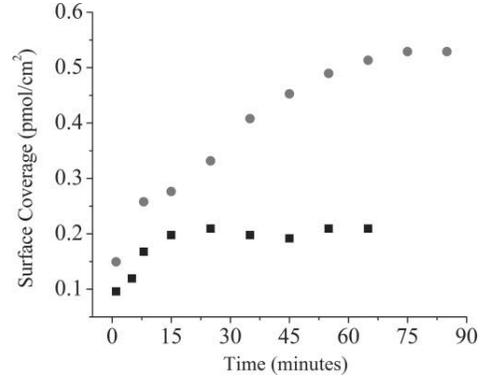


Fig. 9 Adsorption kinetics for chlorophyll a adsorbed onto the hydrophilic surface. In both cases, the amount of chlorophyll a dissolved in solution was the same ($1.4\ \mu\text{M}$), as was the pH of the 7 mM sodium phosphate buffer (7.25). Black squares show that the concentration of NaCl was 1 mM; gray circles show that the concentration of NaCl was 10 mM.

phyll a versus time) onto the hydrophilic surface, and in Fig. 10 the corresponding results for the hydrophobic surface are shown. First, it is worth mentioning that the extremely high sensitivity of our single-mode integrated optical waveguide spectrometer allowed us to investigate surface concentrations as low as a few femtomoles-per- cm^2 . For the hydrophilic surface, the adsorption rate at initial times seems to be independent of the salt concentration, but for a higher salt concentration it takes longer to stabilize and it does so at a higher level of surface coverage. On the other hand, for the hydrophobic surface, the stabilization time seems to be independent of salt concentration, but a higher salt concentration drives a faster adsorption rate at initial times and a larger surface density of chlorophyll a when equilibrium is reached.

The experimental results of surface coverage for chlorophyll a adsorbed to different surfaces and dissolved in different buffer compositions are presented in Table 1; data of surface coverage correspond to values when the equilibrium has been reached. For a $1.4\ \mu\text{M}$ solution of chlorophyll a dissolved in 7 mM of phosphate buffer, we obtained a surface density of $0.209\ \text{pmol}/\text{cm}^2$ for chlorophyll a adsorbed

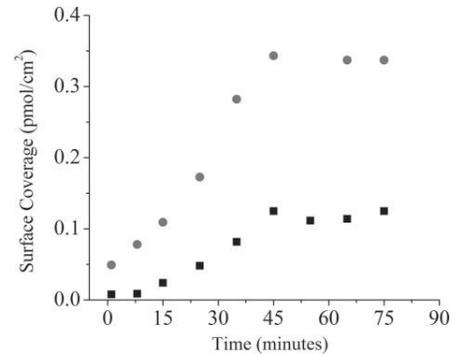


Fig. 10 Adsorption kinetics for chlorophyll a in contact to the hydrophobic surface. In both cases, the amount of chlorophyll a dissolved in solution was the same ($1.4\ \mu\text{M}$), as was the pH of the 7 mM sodium phosphate buffer (7.25). Black squares show that the concentration of NaCl was 1 mM; gray circles show that the concentration of NaCl was 10 mM.

Table 1 Surface density for chlorophyll *a* adsorbed to different surfaces that was equally dissolved in solutions of phosphate buffer with different concentrations of sodium chloride.

Surface Composition	Surface Density for different buffer composition	
	7 mM sodium phosphate plus 1 mM sodium chloride	7 mM sodium phosphate plus 10 mM sodium chloride
178 nm aluminum oxide plus a monolayer of PEG Silane (hydrophilic)	0.209 pmol/cm ²	0.528 pmol/cm ²
178 nm aluminum oxide plus a monolayer of perfluorodecylsilane (hydrophobic)	0.125 pmol/cm ²	0.337 pmol/cm ²

to the hydrophilic PEG surface when the sodium chloride concentration was 1 mM. When the sodium chloride concentration was raised to 10 mM, the surface density increased to 0.528 pmol/cm². For a surface functionalized with a relatively more hydrophobic organic monolayer, the surface density of chlorophyll *a* adsorbed was 0.125 pmol/cm² for a sodium chloride concentration of 1 mM. When the concentration of sodium chloride was raised to 10 mM, the surface density of chlorophyll *a* increased to 0.337 pmol/cm². Therefore, our data show that increasing the concentration of sodium chloride in the dissolving solution dictates a higher surface density of chlorophyll *a*, regardless of the hydrophobicity of the surface. In addition, our data also shows that a higher surface density is reached for a hydrophilic surface indicating that chlorophyll *a* has a higher affinity to hydrophilic surfaces.

3.4 Molar Absorptivity of Chlorophyll *a* Adsorbed to Different Functionalized Surfaces

The molar absorptivity is generally a characteristic spectral fingerprint of a particular material mainly determined by the molecular structure. However, the nano-environment surrounding a particular molecule may affect its structures with possible consequences on the spectral profile of the molar absorptivity. For example, if a molecule is adsorbed to

a surface with a specific chemical composition and physical structure, the surface may perturb the spectroscopic properties of the pristine molecule. In addition, the composition of dissolved species in the solvent may also affect the optical properties of a particular analyte. To obtain the molar absorptivity for the adsorbed chlorophyll *a* across the entire spectral region available from our experimental data, we employed Eq. (1) and the results obtained in Secs. 3.2 and 3.3 for the waveguide sensitivity and surface density for each experiment. The results are presented in Fig. 11 for the hydrophilic surface and in Fig. 12 for the hydrophobic surface. In both Figs. 11 and 12 we have included the molar absorptivity of the dissolved species as obtained from data shown in Fig. 5. As previously mentioned, Fig. 5 clearly indicates that the increase of sodium chloride concentration in the solution from 1 to 10 mM does not affect the molar absorptivity of chlorophyll *a* in the dissolved phase. However, the results reported in Figs. 11 and 12 clearly demonstrate the dependence of chlorophyll *a* upon adsorption on both types of surfaces studied here. For the adsorbed molecules it generated a red-shift (5 nm on the hydrophobic surface and 15 nm on the hydrophilic surface) of the Soret band located at approximately 435 nm. A possible explanation for the observed red-shift of the Soret band at higher ionic strength of the bulk solvent is an aggregation of the chlorophyll *a* molecules. Red-shifted Soret bands have previously been observed for chlorophyll

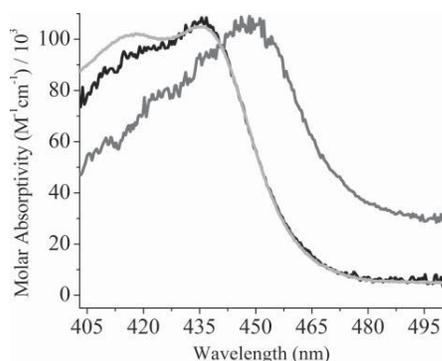


Fig. 11 Molar absorptivity of chlorophyll *a* adsorbed to hydrophilic surface. Black curve: surface-adsorbed chlorophyll *a* that was previously dissolved in 7 mM sodium phosphate buffer (pH 7.25), 1 mM sodium chloride; dark gray curve: surface-adsorbed chlorophyll *a* that was previously dissolved in 7 mM sodium phosphate buffer (pH 7.25), 10 mM sodium chloride; light gray curve: chlorophyll *a* in solution under the same conditions (obtained from a cuvette in a traditional spectrophotometer).

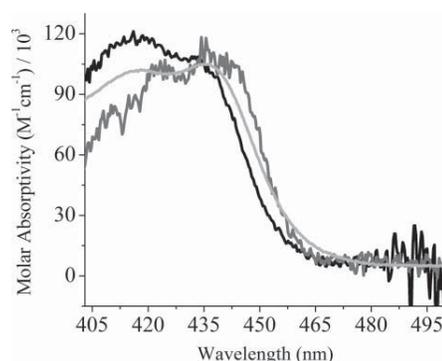


Fig. 12 Molar absorptivity of chlorophyll *a* adsorbed to hydrophobic surface. Black curve: surface-adsorbed chlorophyll *a* that was previously dissolved in 7 mM sodium phosphate buffer (pH 7.25), 1 mM sodium chloride; dark gray curve: surface-adsorbed chlorophyll *a* that was previously dissolved in 7 mM sodium phosphate buffer (pH 7.25), 10 mM sodium chloride; light gray curve: chlorophyll *a* in solution under the same conditions (obtained from a cuvette in a traditional spectrophotometer).

a dissolved in benzene and distributed as a thin liquid film upon a buffered aqueous subphase.²¹ Highly similar shifts of the Soret band have also been observed for chlorophyll dissolved in solutions of water-miscible organic solvents upon the addition of water to the organic solution.^{3,41} A wealth of experimental evidence supports aggregation of the pigment molecules in these solution studies.³ Aggregation in solutions have been attributed to a stronger hydrophobic effect between chlorophyll *a* molecules as the dielectric constant of the solvent is increased. Higher order aggregation of micellar chlorophyll *a* to polymeric forms has also been observed upon the addition of salt (KCl), accompanied by further redshifts in the absorption peaks.³ If the organic layer on the functionalized waveguide surfaces used in this study are behaving as a thin-film of liquid organic solvent, effectively dissolving both some water and chlorophyll *a* (or at least the phytol chain in the case of the hydrophobic surface), then an increase in the ionic strength of the buffer, which increases the polarity of the bulk solvent, might be expected to induce aggregation.

4 Summary and Conclusions

In this work we presented experimental results for the surface coverage and molar absorptivity of chlorophyll *a* adsorbed to hydrophilic and hydrophobic surfaces. This investigation was made possible by our single-mode, broadband, integrated optical waveguide spectrometer, which provides a highly sensitive tool that is able to measure absorbance at submonolayer levels of chlorophyll *a* in contact to a functionalized waveguide surface. Our experiments described here show that the surface composition and sodium chloride concentration in the dissolving solution play an important role in the surface coverage. In both surfaces we observed adsorption of chlorophyll *a*, however for the hydrophilic surface the surface coverage was higher. By increasing the concentration of sodium chloride in the dissolving solution we observed an increase in the absorbance and a red shift in the Soret peak for the hydrophilic surface-adsorbed chlorophyll *a*. The study of adsorption of chlorophyll *a* onto different surfaces presented herein can facilitate the design and understanding of emerging technologies inspired by biomolecular chromophores.

Acknowledgments

The authors acknowledge support from the National Institute of Health (NIH Grant No. RR022864) and National Science Foundation (NSF EPSCoR Grant No. 0814194) both awarded to SBM. We also thank Donald Yeager, Mark Crain, and Professor Kevin Walsh in the Micro/Nano Technology Center of University of Louisville for their support with the atomic and molecular layer deposition tools.

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